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Acids taste sour, while bases taste bitter. Common laboratory acids include hydrochloric acid (HCl), sulfuric acid ( $\text{H}_2\text{SO}_4$ ) and nitric acid ( $\text{HNO}_3$ ). Common laboratory bases include potassium hydroxide (KOH), magnesium oxide (MgO), calcium carbonate ( $\text{CaCO}_3$ ) and sodium



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hydrogen carbonate ( $\text{NaHCO}_3$ ). A strong acid or base completely ionises in ...

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. ... His acid-base theory is sometimes  
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and bases. According to the Lewis theory, an acid is an electron pair acceptor, while a base is an electron pair donor ...

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Monoprotic acid = acid that releases

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one H<sup>+</sup> ion per molecule e.g. HCl  
(hydrochloric acid), HNO<sub>3</sub> (nitric  
acid), CH<sub>3</sub>COOH (ethanoic acid)

Diprotic acid = acid that releases two  
H<sup>+</sup> ions per molecule e.g. H<sub>2</sub>SO<sub>4</sub>  
(sulfuric acid), H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> (ethanedioic  
acid) Moles of acid Moles of H<sup>+</sup>  
Moles of acid Moles of H<sup>+</sup> 3 moles of

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HNO<sub>3</sub> 0.1 moles of H<sub>2</sub>SO<sub>4</sub>

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Chemistry 12 Notes on Unit 4 –

Acids, Bases and Salts Chemistry 12

-Unit 4 -Notes Page 1 Chem 12- Notes

on Acids & Bases After Tutorial 14 &

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Workbook Questions 10-19 p.  
115-121 Do demo of Conductivity 1  
M HCl, 1 M CH<sub>3</sub>COOH Strong &  
Weak Acids & Bases Strong Acid- An  
acid which is 100% ionized in a water  
solution.

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## Salts

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## Answers Acid Bases

TASK 1 – Bronsted-Lowry acids & bases  
1 Acid =  $\text{H}_2\text{O}$ , base =  $\text{NH}_3$   
2 Acid =  $\text{HCl}$ , base =  $\text{H}_2\text{O}$   
3 Acid =  $\text{HCOOH}$ , base =  $\text{KOH}$   
4 Acid =  $\text{HCl}$ , base =  $\text{CH}_3\text{COOH}$   
5 Acid =  $\text{HCl}$ , base =  $\text{NH}_3$   
6 Acid =  $\text{HCO}_3^-$ , base =  $\text{OH}^-$   
7 Acid =  $\text{H}^+$ , base =  $\text{HCO}_3^-$   
8 Acid =

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H<sub>2</sub>SO<sub>4</sub>, base = HNO<sub>3</sub> TASK 2 – pH of  
strong acids 1 a 0.70 b 1.30 c 1.10 d  
-0.30

Chemsheets A2 009 (Acids & bases)  
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Answers

Grade 10 Academic Chemistry

Student Learning Goals/Checklist By

the end of this strand, I will be able to:

use appropriate terminology  
related to chemistry (see vocabulary  
checklist). describe the  
relationships between chemical



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formulae, composition and names of binary compounds.

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Acid Base Study Guide - Multiple  
Choice - Page 1 of 47 Multiple Choice  
Section: This study guide is a  
compilation of questions from

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ACID BASE STUDY GUIDE  
Collins CSEC® Chemistry Workbook  
answers A1 States of matter 1. a) i)

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Ammonium chloride (1) ii) Diffusion  
Diffusion is the movement of particles from an area of higher concentration to an area of lower concentration until the particles are evenly distributed.  
(2) iii) The ammonia solution gave off ammonia gas and

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