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the Periodic Table If
the statement is true,
write true. If the
statement is false,
change the
underlined word or
words to make the
statement true. 1. An
atom ' s valence
electrons are those
electrons that have

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the highest energy. 2.

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Atoms tend to be

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stable and

Answers
nonreactive if they

have six valence

electrons. 3.

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atomic structure &

periodic trends ...

Chapter 5: Atoms,

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Valence electrons

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Periodic Table
and bonding Valence
electrons – electrons

in the highest energy
level (outermost
electron shell) and

are held most loosely
The number of

valence electrons in
an atom determines
its chemical and
physical properties

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Answers

the sharing

of electron pairs

between atoms. The

stable balance of

attractive and

repulsive forces

between atoms when

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Atoms bonding and

breaking, caught on

camera in real-time

for the first time. As

the video shows, the

two atoms bond and

walk together along

the narrow hallways

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of the carbon test
tubes.

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...

The smallest possible amount of matter, which still retains its identity as a chemical element and consists of nucleus surrounded by electrons, is called an atom. It is incapable

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of further division.

These elements are

then grouped in a

periodic table

depending on their

sizes. Take the test

and see what you

know about the

elements.

Watch the first-ever

video of individual

atoms bonding and ...

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Answers

Bonding 1. Handout:

Periodic Table and

Bonding Notes 2.

Periodic Properties

and the Development

of the Periodic Table

1. The first periodic

table was arranged

by Dimitri Mendeleev

in 1869. 1. He was a

professor of

Chemistry. at the

University of St.

Petersburg in Russia

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and was

Periodic Table

Atoms, Bonding, and

the Periodic Table

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Chapter 5 Atoms and

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Table As the atomic

number increases,

the number of

electrons also

increases. As a result,

the properties of the

elements change in a

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regular way across a
period. Chapter 5

Atoms and Bonding

End of Section:

Atoms, Bonding, and
the Periodic Table.

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Periodic trends affect
bonding, because of
how the elements are

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periodic table. For

example elements

can be arranged by

their electronegative,

electron affinity,

atomic radius, or

ionization energy.

Electronegative is the

atoms ability to

attract other bonded

atoms. Electron

affinity is an atoms

ability to attract

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another atom.

Periodic Table

22.1: Periodic Trends

in Bonding -

Chemistry LibreTexts

7 Name two events

involving electrons

that can result in the

formation of

chemical bonds

between atoms. 10

How can you tell

whether or not an

element will react

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Answers

with other elements?
11 What do atoms of elements in the same group of the periodic table have in common?
12 When metals react with other elements, the atoms of the metals _____ electron.

Chapter 5: Atoms,
Bonding, and the
Periodic Table

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Answers

No electronegativity difference between two atoms leads to a

pure non-polar covalent bond. A

small

electronegativity

difference leads to a polar covalent bond.

A large

electronegativity

difference leads to an ionic bond. Polar

bonds and polar

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molecules. In a

Periodic Table

simple molecule like

Answers
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HCl, if the bond is

polar, so also is the

whole molecule.

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atoms elements

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organizes the known
elements of our
world. Check out our
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and properties and
periodic trends. Our
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a free, world-class
education to anyone,
anywhere.

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and ...

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Answers

Periodic arrangement and trends: The columns of the periodic table, which contain elements that show a family resemblance, are called groups. All members of a particular group have analogous outermost (valence) electron configurations, suggesting that all

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members of a group should show a family relationship in the types and numbers of the chemical bonds ...

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