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*Chemistry Solutions Chapter 11 ...*

**NCERT TEXTBOOK QUESTIONS**

**SOLVED. 3.1. How would you determine the standard electrode potential of the system  $Mg^{2+} | Mg$ ? Ans: A cell will be set up**



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*consisting of Mg/MgSO<sub>4</sub> (1 M) as one electrode and standard hydrogen electrode Pt, H<sub>2</sub> (1 atm) / H<sup>+</sup> (1 M) as second electrode, measure the EMF of the cell and also note the direction of deflection in the voltmeter.*

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*Chlorine (Cl) - Structure, Properties, Uses, and FAQs  
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*students who are preparing for the Class 12 chemistry examination and graduate entrance examinations.*

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