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process in which a chemical reaction takes place at the expense of (a) chemical energy (b) electrical energy (c) heat energy (d) none of these

Q.2 Standard hydrogen electrode has an arbitrarily fixed potential (a) 0.00

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volt (b) 1.00 ...

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10) What is the coefficient of Fe^{3+} when the following equation is balanced? ...

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take to plate out 2.19 g of chromium metal from a solution of Cr^{3+} using a ...

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questions. (i) The cathode
reaction in the electrolysis

of dil. H_2SO_4 with Pt

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electrode is (a) Reduction
(b) Oxidation (c) Both
oxidation and reduction (d)
Neither oxidation nor
reduction (ii) Which of the
following statement is
correct about galvanic cell?

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Electrolysis is the process
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takes place at the expense of (a) chemical energy (b) electrical energy (c) heat energy (d) none of these Q.2 Standard hydrogen electrode has an arbitrarily fixed potential (a) 0.00 volt (b) 1.00 volt (c) 0.10 volt (d)

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none of these

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examples?

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Electrolysis of dilute
aqueous NaCl solution was
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milliampere current. The time required to liberate 0.01 mol of H_2 gas at the cathode is (1 Faraday = 96500 C mol⁻¹) [IIT JEE 2008] A. ... ?

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