

## Acces PDF Chapter 111 Describing Chemical Reactions

# Chapter 111 Describing Chemical Reactions

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A chemical equation that does not indicate the relative amounts of the reactants and products Catalyst Substance that speeds the reaction but is not used up in the reaction.

## SECTION 11.1 DESCRIBING CHEMICAL REACTIONS

(pages 321–329)

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Chapter 11 Describing chemical reactions. A representation of a chemical reaction A chemical Equation that does not indicate the relative amount... A substance that speeds up the reaction

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but is not used up in... Small whole numbers that are placed in front of formulas in an... Chemical Equation A representation of a chemical reaction Skeleton Equation A...

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Although the chemical reactions involved in photosynthesis and digestion are different, both chemical reactions are necessary to sustain life. In a chemical reaction, one or more substances (the reactants) change into one or more new substances (the products). Figure 11.1a shows the ingredients for making leavened bread—flour, salt, yeast, and water.

Chemistry chapter 11 chemical reactions packet answers  
Chapter 11 Chemical Reactions 113 SECTION 11.1

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## DESCRIBING CHEMICAL REACTIONS (pages 321–329)

This section explains how to write equations describing chemical reactions using appropriate symbols. It also describes how to write balanced chemical equations when given the names or formulas of the reactants and products in a chemical reaction.

### 11.1 Describing Chemical Reactions 11

Every minute of the day, chemical reactions take place—both inside you and around you. •After a meal, a series of chemical reactions take place as your body digests food. •Plants use sunlight to drive the photosynthetic processes needed to produce plant growth.

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### Chapter 8.1 : Describing Chemical Reactions

DNA is the if all organisms. The pyrimidine bases have a i The  
purine bases have adm e/ ( 3 structure. Solutions in Pearson  
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Chapter 11

Chapter 8.1 : Describing Chemical Reactions 1. Chemical equations and reactions  
Chapter 8.1 2.

Objectives:  
List three observations that suggest that a



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chemical reaction has taken place.<br />List three requirements for a correctly written chemical equation.<br />Write a word equation and a formula equation for a given chemical reaction.<br />Balance a formula equation by ...

### 11.1 Describing Chemical Reactions

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Chapter 11 - Chemical Reactions - 11.1 Describing Chemical Reactions - 11.1 Lesson Check - Page 354: 7 Answer Write the chemical formulas of the reactants on the left of the yield sign and the chemical formulas of the products on the right of the yield sign.

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Section 11.1 Describing Chemical Reactions 323  
withChemASAP Chemical Equations Word equations adequately describe chemical reactions, but they are cumbersome. It's easier to use the formulas for the reactants and products to write chemical equations. A chemical equation is a representation of a chemical reaction; the formulas of the reactants (on

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## 11.2 Types of Chemical Reactions>

Figure 11.1.1: The coffee percolator on the left has been tarnished from exposure to sulfur. Tarnish is the chemical compound silver sulfide. The same percolator on the right has been polished with a tarnish removal product in order to restore its silver finish. (CC BY-NC; CK-12) Another common chemical reaction is the burning of methane gas.

## Chapter 111 Describing Chemical Reactions

a representation of a chemical reaction; the formulas of the reactants (on the left) are connected by an arrow with the formulas of the products (on the right). Write the formulas of

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the reactants to the left of the yields sign (arrow) and the formulas of the products to the right. first write the skeleton equation.

### 11.1 Describing Chemical Reactions 11

Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride.

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11.2 Types of Chemical Reactions> 29 Whether one metal will displace another metal from a compound depends upon the relative reactivities of the

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