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Chapter 12.2: Stoichiometry of Reactions in Solution ...
Chapter 12 Stoichiometry . In the reaction represented by the equation $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$, how many grams of ... Answer the questions above, assuming we started with 30 grams of ammonium nitrate and 50 grams of sodium phosphate. Consider the following reaction:

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Chapter 12 Stoichiometry 12.1 The Arithmetic of Equations 12.2 Chemical Calculations 12.3 Limiting Reagent and ... Sample Problem 12.1 When using conversion factors, remember to ... excess of 1800 is a logical answer. • The unit of the known (FSW 3 HP 2) cancels.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS
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An expanded version of the flowchart for stoichiometric calculations illustrated in Figure 11.4.1 is shown in Figure 12.2.1. We can use the balanced chemical equation for the reaction and either the masses of solid reactants and products or the volumes of solutions of reactants and products to determine the amounts of other species, as ...

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Chapter 11 Small-Scale Lab Section 11.3 Precipitation Reactions: Formation of Solids, page 345 Analysis 1. a. $\text{Na}_2\text{CO}_3 + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{CO}_3(\text{s})$ b. $2\text{Na}_3\text{PO}_4 + 3\text{Pb}(\text{NO}_3)_2 \rightarrow 6\text{NaNO}_3 + \text{Pb}_3(\text{PO}_4)_2(\text{s})$

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Chapter 12 Stoichiometry 127 SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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