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grams problems:  
Using the following  
equation:  $2 \text{NaOH} +$

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reaction represented  
by the equation  $2\text{Na}$   
 $2\text{H}_2\text{O} \rightarrow 2\text{NaOH} +$   
 $\text{H}_2$ , how many grams  
of

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problems; one for  
each reagent given. If  
2 products are given,  
pick one and use it  
for both calculations

If 10.6 g of copper  
reacts with 3.83 g  
sulfur, how many  
grams of the product  
(copper (I) sulfide)  
will be formed?

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Practice: 1. In the  
reaction:  $4\text{Li (s)} + \text{O}_2$   
 $(\text{g}) \rightarrow 2\text{Li}_2\text{O (s)}$  a.

what is the product?

b. what are the

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reactants? c. what

does the “(s)” after  
the formula of lithium  
oxide signify?

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quantities in chemical reactions is a subject of chemistry. Mole ratio. A conversion factor derived from coefficients of a balanced chemical equation interpreted in terms of moles.

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An expanded version of the flowchart for stoichiometric calculations illustrated in Figure 11.4.1 is shown in Figure 12.2.1. We can use the balanced chemical equation for the reaction and either the masses of solid reactants and products or the volumes of solutions

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of reactants and products to determine the amounts of other species, as ...

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