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Chemistry 102 ANSWER KEY REVIEW QUESTIONS Chapter 14

1. A mixture of 0.10 mol of NO, 0.050 mol of H₂ and 0.10 mol of H₂O is placed in a 1.0L flask and allowed to reach equilibrium as shown below:

$$\text{N}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g}) \rightleftharpoons 2\text{NO}(\text{g}) + 2\text{H}_2(\text{g})$$

At equilibrium [NO] = 0.062 M.
Calculate the equilibrium constant, K_c, for this reaction.

Chemical Kinetics Rate Laws – Chemistry Review – Order of Reaction & Equations

Chapter 12 Review Stoichiometry Important Vocabulary •

Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry • Mole ratio- a

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conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles

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Solutions SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Describe the errors made by the following students in making molar solutions. a. James needs a 0.600 M solution of KCl. He measures out 0.600 g of KCl and adds 1 L of water to the solid.

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