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OF EQUATIONS (pages 353–358)

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calculate the amount of
reactants required or
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*nonchemical process. It
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LESSON PLAN 12.4 Objectives

- Calculate the theoretical yield of a chemical reaction from data.*
 - Determine the percent yield for a chemical reaction.*
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*meant by the use of the term
“quantitative’ in the
definition of stoichiometry
on page 354? 2. How does
conservation of mass have to
do with stoichiometry? 3.*

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apply to both chemistry and everyday situations. 12.1.2 Interpret balanced chemical equations in terms of moles, representative particles, mass, and gas volume at STP. 12.1.3 Identify the quantities that are always

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8. The coefficients of a balanced chemical equation tell you the relative number of moles of _____ and _____ in a chemical reaction. 9.

Why is the relative number of moles of reactants and

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