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12.3 Limiting Reagent and Percent Yield >

Chapter 11 Small-Scale Lab Section 11.3 Precipitation Reactions: Formation of Solids, page 345 Analysis 1. a. $\text{Na}_2\text{CO}_3 + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + 3\text{Ag}_2\text{CO}_3(\text{s})$ b. $2\text{Na}_3\text{PO}_4 + 3\text{Pb}(\text{NO}_3)_2 \rightarrow 6\text{NaNO}_3 + \text{Pb}_3(\text{PO}_4)_2(\text{s})$

Chapter 12

12.2 Chemical Calculations > 13 Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved. Mass-Mass Calculations In the laboratory, the amount of ...

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

12.2 Chemical Calculations > Other Stoichiometric Calculations In a typical stoichiometric problem: • The given quantity is first converted toThe given quantity is first converted to moles. • Then the mole ratio from the balancedThen, the mole ratio from the balanced equation is used to calculate the number

12.2 Chemical Calculations >

12.3 Limiting Reagent and Percent Yield > Percent Yield Many factors cause percent yields to be less than 100%. • Reactions do not always go to completion; when a reaction is incomplete, less than the calculated amount of product is formed. • Impure reactants and competing side reactions may

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balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate? 3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate? 4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water.

05 CTR ch12 7/9/04 3:34 PM Page 289 THE ARITHMETIC OF ...

Chapter 12 Stoichiometry 127 SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Chapter 12 Stoichiometry 293 Name ____ Date ____ Class ____ LIMITING REAGENT AND PERCENT YIELD 12.3

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