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Chapter 11 - Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between ...

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Please note that this pace is based on
completing selected sections of the text in
90 classes, approximately 90 minutes each.
Refer to the Course Planning Guide on page v
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11.1 What is stoichiometry? ... 12. 13. 14.
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What are the reactants? C) What are the
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CHAPTER Section 11.1 continued In your textbook, read about mole ratios. Answer the questions about the following chemical reaction. sodium + iron(III) oxide \rightarrow sodium oxide + iron

$$6\text{Na}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow 3\text{Na}_2\text{O}(s) + 2\text{Fe}(s)$$

15. What is a mole ratio? ... Study Guide Chemistry: Matter and Change Chapter 11 ...

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CHAPTER 12 Study ...

Chapter 11 study guide True/False Indicate whether the statement is true or false. ____

1. The actual yield is always lower than the theoretical yield. Matching Match each item with the correct statement below. ? a. stoichiometry ? ? b. mole ratio ? ? c. stoichiometric equivalent ? ____ 2.

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... Chapter 12 Stoichiometry . In the reaction represented by the equation $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$, how many grams of hydrogen are produced if 120. g of Na and 80.0 g of H₂O are available? . For the reaction represented by the equation

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