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Chapter 12 REVIEW: Stoichiometry,
Theoretical, Actual & Percent yield Part I.
Stoichiometry 1. $1 \text{ ZnI}_2 \rightarrow 1 \text{ Zn} + 1 \text{ I}_2$ How
many grams of iodine will you produce if you
begin your rxn with 56.7g of zinc iodide?

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STOICHIOMETRY 12 Vocabulary Review Match the correct vocabulary term to each numbered statement. Write the letter Of the correct term on. the line. Column A 1. the starting materials in a chemical reaction 2. a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms Of moles

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stoichiometry Part A Completion Use this

completion exercise to check your

understanding of the concepts and terms ...

SECTION 12.1 THE ARITHMETIC OF EQUATIONS Use

the 3-step problem-solving approach you

learned in Chapter 1. 1.

Chapter 11 Small-Scale Lab

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chapter-12-solutions-section-2-answers
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chapter-12-stoichiometry-core-teaching ...

05 CTR ch12 7/9/04 3:34 PM Page 289 THE
ARITHMETIC OF ...

Chapter 10 - moles (handouts) Chapter 11 -

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reactions (handouts) Chapter 12 -
stoichiometry (handouts) Chapter 13 - states
of matter (handouts) Chapter 17 -
thermochemistry (handouts) Chapter 18 -
reaction rates (handouts) Chapter 19 - acids,
bases, and salts (handouts) Material Science
Schedule. Previous weeks schedule; Chemistry
Basics ...

chapter-12-section-3-the-business-of-america-
answer-key ...

Chapter 12 (pages 320 – 321) – Answer
questions – 1-9, 14. End-of-Chemistry
Assessment Review. To help you prepare for

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the End-of-Chemistry Assessment, review the following at the end of each chapter: Chapter Highlights. Vocabulary . Answer Chapter Review questions (answers will be posted on my website next week)

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Chapter 12 Review Stoichiometry Important Vocabulary • Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry • Mole ratio- a conversion factor derived from the coefficients of a balanced chemical equation

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interpreted in terms of moles

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Chapter 11 Small-Scale Lab Section 11.3
Precipitation Reactions: Formation of Solids,

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page 345 Analysis 1. a. $\text{Na}_2\text{CO}_3 + 2\text{AgNO}_3 \rightarrow ?$
 $2\text{NaNO}_3 + 3\text{Ag}_2\text{CO}_3(\text{s})$ b. $2\text{Na}_3\text{PO}_4 + 3\text{Pb}(\text{NO}_3)_2$
 $\rightarrow 6\text{NaNO}_3 + \text{Pb}_3(\text{PO}_4)_2(\text{s})$ 2.

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