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reactions in solution are carried out in the same manner as we discussed in Chapter 11. Instead of masses, however, we use volumes of solutions of known concentration to determine the number of moles of reactants. Whether we are dealing with volumes of solutions of reactants or masses of reactants, the coefficients in the balanced chemical equation tell us ...

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units and chemical formulas. 1. Balance the following equation and work the

problems below: $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$

3. If 6.52 L of H_2 react with excess nitrogen what volume of NH_3 is produced?

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1. The first step in most stoichiometry
problems is to _____. a. add the
coefficients of the reagents c. convert
given quantities to volumes b. convert
given quantities to moles d. convert given
quantities to masses _____ 2.

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numbers of molecules or moles. We will first present the method used in most other books for converting from the mass of any reactant or product to the mass of any other reactant or product using a balanced chemical equation outlined ...

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