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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

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12.1: Everyday Stoichiometry - Chemistry LibreTexts

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Chapter 12 ...

Chapter 12 Stoichiometry Review Worksheet Answer Key

Chapter 12 REVIEW: Stoichiometry, Theoretical, Actual & Percent yield Part I. Stoichiometry 1. $1 \text{ ZnI}_2 \rightarrow 1 \text{ Zn} + 1 \text{ I}_2$ How many grams of iodine will you produce if you begin your rxn with 56.7g of zinc iodide? 56.7 g 1 mole 1 mole 253.8 g 319.19 g 1 mole 1 mole = 45.08 grams of Iodine 2. $1 \text{ Pb}_2(\text{CO}_3)_3 + 6 \text{ K} \rightarrow 2 \text{ Pb} + 3 \text{ K}_2\text{CO}_3$

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