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Chapter 13: Electrons in Atoms

Chapter 10 - States of Matter; Chapter 11 - Thermochemistry; Chapter 12 - The Behavior of Gases; Chapter 13 - Electrons in Atoms; Chapter 14 - Chemical Periodicity; Chapter 15 - Ionic Bonding and Ionic

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Compounds; Chapter 16 -
Covalent Bonding; Chapter 17
- Water and Aqueous Systems;
Chapter 18 - Solutions;
Chapter 19 - Reaction Rates
& Equilibrium

Section 13.1 Chapter 13

Electrons in Atoms z

Chapter 13- Electrons in
Atoms ★ Atomic model
evolution: John Dalton- atom
is solid indivisible mass,
explains the nature of
chemical reactions,
discovery of subatomic
particles JJ Thomson-
discovered the electron →
plum pudding model →
negatively charged electrons
stuck to positive mass → no
info on protons and

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electrons or their location
Ernest Rutherford- dense
nucleus, rest empty ...

chapters 13&14 - Chapter 13 Electrons in Atoms ...

Chapter 13 & 14 Assignment & Problem Set 7. An atom of an element has two electrons in the first energy level and five electrons in the second energy level. Write the electron configuration for this atom and name the element.

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Section 13.1 Models of the

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Atom. zOBJECTIVES: Summarize the development of atomic theory. Explain the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

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Electrons in Atoms & Periodic Relationships 4 Chapter 13-14 Assignment & Problem Set •Read Chapter 13 & 14, except skip "Light and Atomic Spectra" p372-375, "The Quantum Concept and the Photoelectric Effect" pp376-378, and "Quantum Mechanics" pp381-382. Be

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sure not to skip over "An Explanation of Atomic Spectra" pp379-380.

Chapter 13 Electrons In Atoms

Chapter 13: Electrons in Atoms. Light is produced when electrons gain a small, specific amount of energy called a quantum. Once they gain a quantum of energy, they "leap" from their ground state, the lowest most stable state for an electron, to their excited state, a higher, less stable state for an electron.

**CHEMSITRY NOTES - Chapter 13
Electrons in Atoms**

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Chapter 13 Electrons in Atoms. when electrons occupy orbitals of equal energy, one electron enters each orbital until all orbitals contain one electron and all electrons will have parallel spins, must fill all boxes with one arrow first.

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Chapter 13 Electrons in Atoms

Bohr proposed a planetary

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model where electrons orbit the nucleus in an elliptical path much as planets orbit the sun-- earth orbits the sun so fast that it does not crash into the sun. Can only orbit at certain distances.

Chapter 13 Electrons in Atoms - Socorro Independent School ...

CHEMISTRY NOTES - Chapter 13 Electrons in Atoms. Goals : To gain an understanding of : 1. Atoms and their structure. 2. The development of the atomic theory. 3. The quantum mechanical model of the atom.

Chapter 13 Electrons in

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Atoms - Mrs. Morales PEP site

12. The maximum # of electrons allowed in an energy level = $2n^2$ (where n is the level number) Ex:
level 1 $2(1)^2 = 2$ electrons
max. level 2 $2(2)^2 = 8$ electrons max. 13.

Sublevels: s, p, d, f some people don't forget. s sublevel and orbitals. s sublevels have 1 orbital. Each orbital can hold 2 electrons. Spherical shaped orbital

Electrons in Atoms & Periodic Relationships Chapter 13-14 ...

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PPT - Chapter 13 Review

Electrons in Atoms

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Chapter 13 Homework

Quantum Mechanical Model The energy levels are not equally spaced like a ladder - they get closer the farther from the nucleus you

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go The higher the energy of the e^- , the easier it leaves the atom

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Chapter 13 - Electrons in Atoms Chapter 13: 1 - 20, 23 - 25, 27, 31, 32, 34 - 38, 41, 45, 47, 48, 52 Section 13.1 - Models of the Atom Section Review 13.1 1. List in chronological order, a major contribution of each of these scientists to the understanding of the atom: proposed that all elements are composed of atoms. Dalton

PPT - Chapter 13 Review Electrons in Atoms

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PowerPoint ...

Chapter 13: Electrons in Atoms The Evolution of Atomic Models Section 13.1 Describe an atom. What are the three main subatomic particles? Identify the relative electrical charges associated with each particle.

CHAPTER 13 Electrons in Atoms

Chapter 5 Electrons in Atoms . Name Date 11. The number of sublevels in an energy level is equal to the square of the principal quantum number of that energy level. ... 13. As many as four electrons can occupy the same orbital. 14. The Pauli

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exclusion principle states that an atomic orbital may describe at most two electrons.

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Chapter 13: Electrons in Atoms. The Bohr model limits electrons to specific circular paths. The quantum mechanical model expresses the probability of finding an electron in a given location within the electron cloud based on its current energy level.

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Electrons move in circular orbits . around . the

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nucleus . at . fixed.
energy. levels. Electrons
are never between energy
levels or energy shells. An
electron must have . just
the right amount of energy.
to jump from one level to
another. A . quantum. of
energy is . just the right
amount of energy. needed for
an electron to jump levels.

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