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SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471-477) This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves. Solution Formation (pages 471-472)

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Ch. 16 Textbook Answers 1. Henry's Law states that the solubility of a gas in a liquid is directly proportional to the pressure above the liquid. $S_1 = S_2 \frac{p_1}{p_2}$ $0.16 \text{ g/L} = 0.44 \text{ g/L} \frac{p_1}{104 \text{ kPa}}$ $p_2 = 288 \text{ kPa}$ 2. Henry's Law states that the solubility of a gas in a liquid is directly proportional to the pressure above the liquid. $S_1 = S_2 \frac{p_1}{p_2}$ 3.6

$g/L = 9.5 \text{ g/L}$ p 1 p 2 1 atm 2.6 atm 3. The factors that ...

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Manual "Dynamics ", Chapter 13 Solution Manual Dynamics Hibbeler -
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initial angular velocity of 2 and has a constant angular acceleration
of.

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A property of solutions that depends only on the number of solute
particles, not on their identity. Freezing-point depression The
difference in temperature between the freezing point of a solution and
the freezing point of the pure solvent

Chapter 16 Solutions Answers

Chapter 16 Solutions I. Solutions A. Solution is a homogeneous mixture
involving two or more pure substances. Its composition usually can be
varied within certain limits. B. Solute substance dissolved in the
solution. C. Solvent the substance in which the solute is dissolved
Example: Salt + H₂O H₂O is the solvent NaCl Salt is the solute Na+Cl-
II.

Solution to Chapter 16 Homework

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numbered Conceptual Questions can be found in the back of the book A
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