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Chapter 2 The Chemistry of Life Worksheets  
Chapter 2 The Chemistry of Life Section Review 2-1

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Section Review 2-4 1. protons; neutrons 2. electrons 3. neutrons 4. electrons 5. ionic 6. The two main types of chemical bonds are ionic and covalent bonds.

Chapter 2 Chemistry, Matter, and Life Vocab  
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Chapter 2: Biochemistry Problems 7) Draw several molecules on the screen and calculate their formulas by hand. Check your work by clicking the “ Calculate ” button. Note that it is possible to draw molecules that the software cannot process properly. Some of these molecules are chemically possible, but their chemistry is beyond the scope

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Chemistry you need to know Chapter 2 P 26: Know symbols for elements in table 2.1, oxygen to iron; know charge on ions: calcium, potassium, sodium, chlorine P 29 isotope radioisotopes p 29-30 P 30-31 mixtures, solutions, concentration in percent and molarity (in most medical applications, concentration is in percent)

Chapter 2: Biochemistry Problems

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Review 2-1 Section Review 2-4 1 Chapter 2 the  
chemistry of life test answer key. protons; neutrons 2.

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electrons 3. neutrons 4. electrons 5. ionic 6.

Unit 1 The Chemistry of Life: Chapter 2 - Ms. Kots  
Biology ...

1 Chapter 2.1, 2.2 Review Packet – Answer Key 2.1

The Nature of Matter Lesson Summary Atoms The atom is the basic unit of matter, made up of three subatomic particles. Protons have a positive charge and neutrons carry no charge. Strong forces bind protons and neutrons together in the nucleus.

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12. 4. Problem ... CHEMISTRY 12. 6. Problem Rank the following compounds in sequence from lowest melting point to highest melting point, and give reasons for your decisions: AsBr ...

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Unit 1 The Chemistry of Life: Chapter 2. Unit 2  
Ecology: Chapters 3 to 6: Unit 3 Cells & Cellular  
Processes: Chapters 7-10. Unit 4 Genetics: Chapters  
11-15. ... Packet: Friday, October 2, 2015 Chapter 2  
Section 3 ...

Ch. 2 Answer Key - Lawndale High School  
The Chemistry of Life ANSWER KEY Section 2 – 1 The  
Nature of Matter (pages 35 – 39) This section identifies



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the three particles that make up atoms. It also explains how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds. Atoms (page 35) 1. The basic unit of matter is called a(an) atom .

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Testbank General, Organic, Biological Chemistry, 6th  
edition 1 CHAPTER 2: MEASUREMENTS IN  
CHEMISTRY MULTIPLE CHOICE 1) The

“ mathematical meaning ” associated with the metric  
system prefixes centi, milli, and micro is, respectively,  
a)  $10^{-2}$ ,  $10^{-4}$ , and  $10^{-6}$  c)  $10^3$ ,  $10^6$ , and  $10^9$  b)  $10^{-2}$ ,  $10^3$ , and  $10^{-6}$  d)  $10^{-3}$ ,  $10^9$ , and  $10^{-12}$

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Section Review 2-2 1. a 2. b 3. c 4. c 5. a 6. b 7. c 8.

Polarity in a water molecule is caused by an uneven distribution of electrons between the oxygen and hydrogen atoms. 9. The concentration of H ions determines whether a solution is acidic or basic. 10. Capillary action is the effect of water rising in a narrow tube against the force of ...

Chemistry of Life packet 1 ANSWER KEY (1) - The

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Chemistry ...

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Chapter 2: Introduction to the Chemistry of Life Figure 2.1 Foods such as bread, fruit, and cheese are rich sources of biological macromolecules. The elements carbon, hydrogen, nitrogen, oxygen, sulfur, and phosphorus are the key building blocks of the chemicals found in living

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