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ANSWERS I. Determine what is
oxidized and what is reduced
in each reaction. Identify
the oxidizing agent and the
reducing agent, also. 2SrO
 SrO to Sr^{2+} ;
oxidized/reducing agent 00

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to O^{2-} ; reduced/ox. ag. 1.

$2Sr + O_2 \rightarrow 2SrO$ 2. $2Li + S \rightarrow Li_2S$ LiO

to Li^{1+} ; oxidized/red. ag.

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636 Chapter 20 Redox
Reactions Figure 20-1 The
reaction of magnesium and
oxygen involves a transfer

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of electrons from magnesium to oxygen. Therefore, this reaction is an oxidation–reduction reaction. Using the classifications given in Chapter 10, this redox reaction also is classified

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as a combustion reaction. X
X X X X X X X 2Mg 2MgO Each

...

Writing Half Equations
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redox couples. such as Zn /

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Zn^{2+} and Fe^{2+} / Fe^{3+}
(reduced form / oxidized
form) when characterizing
redox reactions. (Recall the
. half-reaction . method of
balancing redox equations).
In this exercise we deal
with . electrochemical

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cells, in which two redox couples are placed in separate compartments as opposed to their direct contact. Cell ...

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Difference • The energy that
would be required to move a
unit of electrical charge
from ... Answer: -0.40 V .
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I. Determine what is

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oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

1. $2\text{Sr} + \text{O}_2$
2. $2\text{Li} + \text{S}$
3. $2\text{Cs} + \text{Br}_2$

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in each reaction. Identify the oxidizing agent and the reducing agent, also.

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reduced in each reaction.

Identify the oxidizing.

agent and the reducing

agent, also. 1. $2\text{Sr} + \text{O}_2$

2SrO Sr 0 to Sr $2+$;

oxidized/reducing agent O 0

to O $2-$; reduced/ox. ag.

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