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AP Chemistry CHAPTER 21-

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Nuclear Chemistry 21.1

Radioactivity •When nuclei change spontaneously, emitting energy, they are said to be radioactive.

•Nuclear chemistry is the study of nuclear reactions and their uses.

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21.2: Patterns of Nuclear
Stability. 21.2.1 Neutron-to-
Proton Ratio. strong nuclear
force - a strong force of

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attraction between a large number of protons in the small volume of the nucleus; stable nuclei with low atomic numbers up to 20 have nearly equal number of neutrons and protons

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nuclide. mass defect.
nuclear binding energy. a
proton or neutron. an atom,
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Flashcard. nucleons. protons
and neutrons. nuclide. An
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number of protons and neutrons in its nucleus. mass defect. The difference between the mass of an atom and the sum of the masses of its protons, neutrons, and electrons.

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below just once to complete
the following sentences.

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reactor beta particle half-
life radioactivity deuterium

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gamma ray nuclear fusion
tritium 1.

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I'll first show you how to determine an element's number of protons, electrons, and neutrons from its atomic symbol. I'll also teach ...

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Chemistry: Part 1 of 9

In this video I'll do an example problem that involves a calculation using Einstein's energy-mass equation: $E = mc^2$.

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Chapter 21: Nuclear
Chemistry 21.1: The Nature
of Nuclear Reactions
Nucleons: - the particles

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that make up a nucleus of an atom (protons, $(1\ 1\ p\ +\ \text{or}\ 1\ 1\ H)$ and neutrons, $(1\ 0\ n)$).
Isotopes: - atoms that have different mass number but the same atomic number or number of protons.

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Chapter 21-Assignment C:
Summary and Review You may
think of nuclear chemistry
as an untamed jungle, but

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there are rules to help you find the trails, just as you found the rules and trails in ordinary chemical reactions.

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Modern Chemistry 171

Nuclearchemistry CHAPTER 21

REVIEW Nuclear Chemistry

SECTION 2 SHORT ANSWER

Answer the following

questions in the space

provided. 1. _____ The

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nuclear equation is an example of an equation that represents (a) alpha emission. (b) beta emission. (c) positron emission. (d) electron capture. 2.

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Nuclear Chemistry Kinetics
of Radioactive Decay A

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wooden object from an archeological site is subjected to radiocarbon dating. The activity of the sample that is due to ^{14}C is measured to be 11.6 disintegrations per second.

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Chemistry. History Just
because it is interesting.
Marie Curie: Henri Becquerel
(Physics professor, 1852 ...

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Nuclear particles involved
in nuclear reactions ...
effect is mass same, protons
up one thus neutrons down
one. Example: ${}_{20}^{42}\text{Ca} \rightarrow {}_{21}^{42}\text{Sc} + {}_{-1}^0\text{e}$; Alpha emission
(occurs when the nucleus is
too large ...

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