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CHAPTER 21 REVIEW

Nuclear Chemistry

SECTION 2 SHORT

ANSWER Answer the

following questions

in the space

provided. 1. _____

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The nuclear equation is an example of an equation that represents (a) alpha emission. (b) beta emission. (c) positron emission. (d) electron capture. 2.

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History Just because
it is interesting. Marie

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Curie: Henri
Becquerel (Physics
professor, 1852 -
1908) discovered
radioactivity in 1896
by placing a
photographic plate
near a box of
uranium ore. When
the plate "developed"
on its own, he knew
the ore contained
something giving off
energy.

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22Ne?? Nuclear
Reactions Unstable
nuclei undergo
spontaneous changes
that change their
num-ber of protons

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and neutrons. In this process, they give off large amounts of energy and increase their stability. These changes are a type of nuclear reaction. A nuclear reaction is a reaction that affects the nucleus of an atom.

Chapter 21 Nuclear Chemistry

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152 Nuclear
chemistry involves
changes in the
nucleus (protons and
neutrons) of
radioactive atoms.

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medical diagnosis
and treatment C-14
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plants create new

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STUDY. PLAY.

22.2 Answers

Radioactive decay.

the spontaneous
disintegration of a
nucleus into a slightly
lighter and more
stable nucleus,
accompanied by
emission of particles,
electromagnetic
radiation, or both ...

CHAPTER 22 Nuclear
Chemistry

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Nuclear Chemistry-
Chapter 22. STUDY.

PLAY. nucleons.
description of atomic
nuclei based on
proton and neutrons
particles. nuclide.
used in nuclear
chemistry as a
reference to an atom,
identified by the
number of protons
and neutrons in its
nucleus (another

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name for an isotope

... Chapter 22 Review:

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terms. common ions

(name ...

CHAPTER 22 REVIEW

Nuclear Chemistry -

pnhs science

Chapter 22 Nuclear

Chemistry 22.1

Multiple Choice

Questions 1) When

more than 3600

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known nuclides are plotted on a neutron/proton grid they make up a group called A) the "island of stability." B) the "peninsula of nuclear stability." C) the "sea of instability."

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Nuclear

Transformations •

Rutherford in 1919

performed the first

nuclear

transformation. •

The transmutations

are sometimes

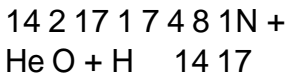
represented by listing

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in order, the target
nucleus, the

bombarding particle,
the ejecting particle
and the product

nucleus. • The above
equation becomes:



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Nuclide a distinct

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22.3 Answers
kind of atom or
nucleus characterized
by a specific number
of protons and
neutrons Mass Defect
difference between
the mass of an
isotope and its mass
number Nuclear
Binding energy is the
energy that would be
required to
disassemble the
nucleus of an atom

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into its component
part which would be
neutrons [...]

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Questions | Nuclear
Physics ...

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Nuclear Chemistry
SECTION 22-4 SHORT
ANSWER Answer the
following questions
in the space
provided. 1. Label

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each of the following
statements with one
of the choices below:
(1) fission only (3)
both fission and
fusion

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