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Section 6 1

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Covalent Bonding.
A covalent bond is
a chemical bond in
which two atoms*

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share a pair of valence electrons. When two atoms share one pair of electrons, the bond is called a single bond.

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2. Covalent

Bonding and

Molecular

Compounds.

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Bonding

STUDY. PLAY.

Diatomic molecule.

a molecule

containing only two

atoms. Anion. a

negative ion.

Cation. a positive

ion. Molecules. a

neutral group of

atoms that are held

together by

covalent bonds.

Molecular

compound.

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Bonds Section 6 1

Answer Key

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Chemical Bonding

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abouturier.

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6 section 3,4,5.

ionic compound.

formula unit. lattice

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energy. polyatomic ion. A compound that consists of positive and negative ions. the lowest whole-number ratio of ions in an ionic compound.

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Modern Chemistry

9 Chemical

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Bonding CHAPTER 6 STUDY GUIDE

Chemical Bonding SECTION 3 IONIC BONDING AND IONIC COMPOUNDS SHORT ANSWER

Answer the following questions in the space provided. 1. _____

The notation for sodium chloride, NaCl, stands for

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Bonding SECTION 3

SHORT ANSWER

*Answer the
following questions
in the space*

*provided 1 a The
notation for sodium
chloride, NaCl,
stands for one (a)
formula unit (c)
crystal (b)
molecule (d) atom*

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Chemical Bonds

2 d In a crystal of
an ionic compound,
each cation is ...

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Section 1 Chemical

Bonding Answers ...

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Chemical Bonding

SECTION 6-2

SHORT ANSWER

Answer the

following questions

in the space

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Chemical Bonds

Section 6.1 Ionic

provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. 2. Name two elements that form compounds that are exceptions to the octet rule.

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Chemical Bonds

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer

the following

questions in the

space provided. 1.

a A chemical bond

between atoms

results from the

attraction between

the valence

electrons and of

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Chemical Bonds

Section 6.1 Ionic Bonding

- different atoms. (a)*
 - nuclei (c) isotopes*
 - (b) inner electrons*
 - (d) Lewis structures*
- 2. b A covalent bond consists of (a) a shared electron.*

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Two or more atoms

held together by a

chemical bond is

called what?

answer choices . A

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AND FORMULAS

Section 6.1

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6.2 REPRESENTING

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COMPOUNDS

Section 6.3 IONIC

CHARGES Section

6.4 IONIC

COMPOUNDS

Section 6.5

MOLECULAR

COMPOUNDS AND

ACIDS Section 6.6

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What is a chemical bond and why does it form?

Section 2 Covalent Bonding and Molecular Compounds

What is a molecular formula? What are the characteristics of a covalent bond? How does a Lewis structure represent

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a molecule?

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Chemical Bonds

SECTION 3 SHORT

ANSWER Answer

the following questions in the space provided.

- 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom.*
- 2. d In a crystal of an ionic compound,*

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Chemical Bonds

Section 6.1 Ionic

each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) positive ions. (d) negative ions.

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Bonds Section 6.4

The Structure Of

Metals

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3 Chemical

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Chemical Bonds

Section 6.4 Ionic

differentiate between covalent and ionic bonding; explain why bonding occurs; use the difference in electronegativity to determine whether a bond is polar covalent, nonpolar covalent or ionic

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Compounds

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Chemical Bonding

SECTION 1 SHORT

ANSWER Answer

the following

questions in the

space provided. 1.

a A chemical bond

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Section 6.1 Ionic

attraction between

the valence

electrons and of

different atoms. (a)

nuclei (c) isotopes

(b) inner electrons

(d) Lewis

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Bonding Section 2

Covalent Answer

Key ...

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in the space*

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SHORT ANSWER

Answer the

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Section 6.1 Ionic

following questions in the space provided.

1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom.

2. d In a crystal of an ionic compound, each cation is ...

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