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## **SPM Form 4 Chemistry**

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**Chapter 7 - Acids  
and Bases**

Chapter 7 -  
Mixtures of Acids  
and Bases  
Introduction In  
Chapter 6, we  
examined the  
equilibrium  
concentrations in  
solutions of acids  
and solutions of  
bases. In this  
chapter, we

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continue our  
discussion of acids  
and bases by  
focusing on the  
equilibrium  
concentrations of  
solutions formed  
by mixing acids  
and bases.

## **Chapter 7 Acids and Bases - Concept Map**

This Page. has



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moved to a new  
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Acids and Base -  
Expected learning  
outcomes.

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a substance that can act as either an acid or base (ex. water). An amino acid has a basic amine group and an acidic carboxylic

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Purple  
acid group, so it is  
also amphoteric.

Unlike water,  
amino acids can  
carry multiple  
charges depending  
on their  
environment. At pH  
7, most amino  
acids have a  
protonated amine  
group and a  
deprotonated  
carboxylic acid

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group.  
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38 Chapter 7 7.15

(HF H<sub>2</sub>SO<sub>4</sub>) + H

2SO<sub>4</sub>(l) ⇌ (H<sub>2</sub>

F+(H<sub>2</sub>SO<sub>4</sub>) +

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$\text{HSO}_4^-$  and  $\text{H}_2\text{SO}_4$  are acting as a base and  $\text{H}_2\text{F}^+$  is the conjugate acid,  $\text{H}_2\text{SO}_4$  is acting as an acid and  $\text{HSO}_4^-$  is the conjugate base. 7.17  $\text{HSeO}_4^-$  (base),  $\text{H}_2\text{SO}_4$  (conjugate acid); H

**Chapter 7 acids  
and bases -  
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## Chapter 7: Neutralization Reactions .

Neutralization  
Reaction. A  
reaction between  
an acid and a base  
to produce a salt  
and water.

(Remember you  
will need to be able  
to write the ionic  
and net ionic  
equations for these

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reactions.)

**Chapter 7:**

Chapter 7 Mixtures  
of Acids and Bases  
181 7.2 BUFFERS In  
Chapter 6, we  
treated in some  
detail four types of  
acid-base  
solutions: strong  
acid, strong base,  
weak acid, and  
weak base. We



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now treat buffers, the last type of acid-base solution to be considered. A buffer is defined in the dictionary as a “device that softens the shock of a blow.”

### **Chapter 7. Acids, Bases, and Equilibrium**

Chapter 7 Acids,

*Page 17/34*

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Properties of Acids  
and Bases Litmus  
is an example of an  
indicator, a  
compound that  
changes color  
when in contact  
with an acid or a  
base.

## **Chapter 7 - Acids and Bases -**

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Chapter 7 Acid-  
Base Equilibria 2.  
Acid-Base Theories

- Arrhenius theory (Nobel Prize) An acid is any substance that ionizes (partially or completely) in water to give hydrogen ions,  $H^+$  (which associate with the solvent to

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give hydronium  
ions,  $H_3O^+$ )  $HA +$   
 $H_2O \rightleftharpoons H_3O^+ + A^-$  •

A base: any  
substance that  
ionizes in water to  
give hydroxyl ions  
 $OH^-$ .

## **Chapter 7 - Mixtures of Acids and Bases**

Chapter 7 Acids,  
Bases, and

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Solutions

Instructions:

Complete the  
crossword puzzle.

Use the clues to  
help identify the  
words. C O N C E N  
T R A T E D H O E I  
S C A L E Y L U L O  
D L T U N S A T U R  
A T E D R O R T A R  
O I A E L O G D L T  
S A T U R A T E D I  
N D I C A T O R I N

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## And Solutions Z V A S S U P E R S A T U R A T E D Puzzle

### **Chapter 7 ACIDS AND BASES**

The p H of a salt solution obtained by the reaction between a strong acid and a strong base is... [A] is 7  
[B] is less than 7  
[C] is more than 7  
[D] cannot be

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determined;

Sodium chloride is  
a salt of... [A] a  
weak acid and a  
weak base [B] a  
strong acid and a  
strong base [C] a  
weak acid and a  
strong base [D] a  
strong acid and a  
weak base

**Chapter 7- Acids  
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**Chapter 7 Acids,  
Bases, and  
Solutions**

(a) Both acids and bases change colour of all indicators. (b) If an indicator gives a colour change with an acid, it does not give a change with a base. (c) If an indicator changes

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Puzzle  
colour with a base,  
it does not change  
colour with an acid.

(d) Change of  
colour in an acid  
and a base  
depends on the  
type of the  
indicator.

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Puzzle

First, according to the Brønsted-Lowry definition of acid-base, acetic acid,  $\text{CH}_3\text{CO}_2\text{H}$ , produces  $\text{H}^+$  thus it is an acid. We can write the equilibrium constant for the acid is, then  $K_{\text{acetic acid}} = \frac{[\text{H}^+][\text{CH}_3\text{CO}_2^-]}{[\text{CH}_3\text{CO}_2\text{H}]}$  =

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$1.74 \times 10^{-5}$  The  
numerator terms  
are the product  
ions

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Chapter 7- Acids  
and Bases. STUDY.  
PLAY. taste sour,  
turns litmus red,  
reacts with active

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And Solutions

Cross Word

metals to form  
hydrogen gas,  
react with bases to  
form water and a  
salt. Characteristics  
of acids. taste  
bitter, turns litmus  
blue, slippery,  
reacts with acids to  
form water and a  
salt.

## **Chapter 7: Acids and Bases**

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navigation

**Chapter 7 Acids,  
Bases, and  
Solutions**

Chapter 7: Acids,

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## And Solutions

Bases, and  
Solutions Solution -  
a homogeneous  
mixture Solutions  
have the same  
properties  
throughout,  
containing solute  
particles  
(molecules or ions)  
that are too small  
to see Solvent -  
the part of the  
solution that is



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present in the  
largest amount and  
dissolves the other  
substances

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35 Chapter 7

ACIDS AND BASES

Exercises 7.2 (a) A  
nonpolar solvent  
because

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phosphorus

pentachloride is a

nonpolar covalent

compound. (b) A

polar protic solvent

because cesium

chloride is ionic.

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