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sharing of electrons occurs the  
attachment between atoms is called.

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covalent bond. In a covalent bond, the dissociation energy is released in the process of an exothermic reaction.

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Single Covalent Bonds When only one

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pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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bond is formed between two atoms when an electron is shared as opposed to lost. Overall, this makes each of the atoms involved more stable. There are specific rules that are used when naming binary molecular compounds, binary acids, and oxyacids.

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Molecular Structure 8-4 H<sub>2</sub> molecule.  
More sophisticated descriptions of  
chemical bonding will be discussed in  
Chapter 9. 8.3 Lewis Structures OWL  
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most important tools chemists use to predict the properties of a chemical species is its Lewis structure.

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Complete Table 8.10 , which shows the  
number of electrons shared in a single  
covalent bond, a double covalent bond,  
and a triple covalent bond. Identify the

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group of atoms that will form each of these bonds.

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and Change • Chapter 8 121 Section 8.1  
The Covalent Bond pages 240–247  
Practice Problems page 244 Draw the  
Lewis structure for each molecule. 1.  
PH<sub>3</sub> H<sub>2</sub> H<sub>2</sub> H— H H P respectively, for  
single, double, and triple P — — 2. H<sub>2</sub> S  
H H H — H S S ...

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Chapter 8 – Covalent Bonding. Jennie L. Borders. Section 8.1 – Molecular Compounds. A covalent bond is formed between atoms held together by sharing electrons. A molecule is a group of atoms joined by covalent bonds. A

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diatomic molecule is 2 atoms bonded together. Diatomic Elements.

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bond and a covalent bond differ in the location of the electrons. In ionic bonds, there is no sharing of electrons. The cation is deficient in electrons and the anion has extra electrons.

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1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases.
- (8) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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Covalent Bonding Or How I Learned to  
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Sharing is Illegal) 2. As you should  
remember, ionic compounds are solids  
at room temperatures that have one ion  
strip the electron(s) from the other

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elements' electron cloud. Some compounds do not give up their electrons so easily.

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Covalent Bond Main Idea: Atoms gain stability when they share electrons and form covalent bonds Why do atoms bond? Non Metal & Metal - Non Metal & Non Metal - Diatomic Elements - Orbital Overlap 1

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*Page 31/33*

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A covalent bond in which one atom contributes both bonding electrons

Polyatomic ion A tightly bound group of atoms that has a positive or negative charge and behaves as a unit.



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