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Bonding and
Molecular Structure
8-2 8.1 Interactions

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Between Particles: Coulomb's Law

OWL Opening Exploration 8.1

Coulomb's Law

Matter is made up of atoms and ions that experience both attractive and repulsive forces.

The strength of the force holding oppositely charged particles together

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Chapter 8

Concepts of

Chemical Bonding

Section 8.4 – Polar

Bonds and

Molecules.

Covalent bonds

involve sharing

electrons between

atoms. When the

atoms in the bond

pull equally, the

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bonding electrons are shared equally, and the bond is nonpolar. When the atoms in the bond pull unequally, the bonding electrons are pulled closer to one atom, and the bond is polar.

Section Vocabulary
- SharpSchool
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BONDING Class 8.2

8.2 8.4 8.3 8.3 8.3

195 Vocabulary

Review Select the

term from the

following list that

best matches each

description.

Chapter Quiz loose

the best answer

and write its letter

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on the line. . A
bond in which each
atom contributes
two electrons is ...

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Chemistry Chapter
8 Covalent Bonding
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Chemistry -

Chapter 8 -

Covalent Bonding.

the octet rule

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cannot be satisfied in molecules whose total number of electrons is an odd number; there are also molecules in which an atom has fewer, or more,

than a complete octet of valence electrons.

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Covalent Bonding
and Molecular

Structure
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Chapter 8 Notes -

Bonding: General

Concepts . 8.1

Types of Chemical

Bonds . A. Ionic

Bonding 1.

Electrons are

transferred 2.

Metals react with

nonmetals 3. Ions

paired have lower

energy (greater

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stability) than
separated ions B.

Coulomb's Law 1.

$$E = -\frac{1}{4\pi\epsilon_0} \frac{Q_1 Q_2}{r}$$

$$2.31 \times 10^{-19} \text{ J}$$

$$\text{nm. } 1.2 \text{ a. } E =$$

energy in joules b.

Q. 1. and . Q. 2

Q Q E r -

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Bonding. 8.1 The

Covalent Bond 8.2

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8.3 Molecular

Structures 8.4

Molecular Shapes

8.5
Electronegativity &
Polarity.

(Chapter 7)

This video explains
the concepts from
your packet on
Chapter 8 (Basic
Concepts of

Page 14/34

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Covalent Bonding),
which can be found
here: [https://goo.gl/
TyuJ36](https://goo.gl/TyuJ36) Section 8...

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242 Chapter 8 □
Covalent Bonding
Single Covalent
Bonds When only
one pair of
electrons is shared,
such as in a

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hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Chem Chapter 8 -
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Review Packet
Flashcards ...

phosphate ion that
has only 8
electrons around
the central
phosphorus, a
common Lewis
structure puts a
double bond
between the

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phosphorus and
one of the oxygens.

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Concepts of

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AM . Name Quarter

Unit 1 Homework

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Covalent Bonding Basics Homework Packet Covalent

Basics 1. What do atoms do with electrons in a covalent bond?

Share them 2. ...

exhibits H-bonding and substance B (density 1.23 g/mL)

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Chemistry Chapter

8- Covalent

Bonding. a chemical bond consisting of a hydrogen atom between two electronegative atoms (e.g., oxygen or nitrogen) with one side be a covalent bond and the other being an ionic bond.

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Chapter 8 Covalent
Bonding Packet

non-polar covalent
bond. a covalent
bond formed by
the equal sharing
of bonding
electrons by two
atoms. hydrogen
bond. force that
occurs when a
hydrogen atom

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that is covalently bonded to a very electronegative atom is also weakly bonded to an unshared pair of electrons in the same or a nearby molecule.

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Chemistry: Matter
and Change

Chapter 8 44 .

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CHAPTER FOR

Class Section 8.2

continued ...

Differentiate

between an

ordinary covalent

bond and a

coordinate

covalent bond.

Give an example of

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a molecule that exhibits both and label them. —each 0.40M Shares

Sharæ 4. Most elements follow the octet rule.

Chemistry Chapter 8 Covalent Bonding Flashcards | Quizlet
In water's two covalent H—O bonds, the

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electrons in the bond are not shared equally by the two atoms.

Oxygen, which has a stronger attraction for electrons than hydrogen, pulls the electrons towards itself. This creates a polar covalent bond -----> _____ sharing. In a polar

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covalent bond, the
more

electronegative
element

Chemistry -

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Flashcards | Quizlet

that are introduced

in this section.

Each blank can be

completed with a

term, short phrase,

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or number The
quantum
mechanical model

of bonding

assumes that 1.

f(C) r rh atomic

orbitals overlap to

produce 1 A

molecular orbit

that 2-can be

occupied by two

electrons of a

covalent bond is

called a 3. S

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Bonding. Valence
shell electron pair

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repulsion theory;
because electron
pairs repel,
molecules adjust
their shapes so
that valence
electron pairs are
as far apart as
possible.

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□ Recall that ionic
bonds form when
the combining

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atoms give up or
accept electrons.

□ Another way that
atoms can combine
is by sharing
electrons.

Molecules and
Molecular

Compounds

Sharing Electrons

-Atoms that are
held together by
sharing electrons
are joined by a

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8.3 Bonding
theories. essential
Understanding
Scientists use a
variety of theories
and models to
explain how and
why covalent

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bonds form. Lesson
summary.

Research
Education

molecular orbitals

One model of
molecular bonding
pictures a
molecular orbital
that. is a
combination of
individual atomic
orbitals. A bonding
orbital can be
occupied by a pair
of electrons.

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