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42. a. 2 – b. 1+ c. 1 – d.

3+ 43. a. 2+ b. 2+ c. 3+

d. 1+ 44. a. barium ion b.

iodide ion c. silver ion d.

mercury(II) ion 45.

cyanide, CN^- and

hydroxide, OH^- 46. a.

hydroxide ion b.

lead(IV) ion c. sulfate ion

d. oxide ion 47. zero; A

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compound is electrically neutral. 48. The symbols for the cation and anion are written and the charges

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MANUAL Problem-

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Halogen Atomic Radius

(ppm) Ionic Energy

(kJ/mol) Electronegati-

vity Fluorine 72 1681

3.98 Chlorine 100 1251

3.16 Bromine 114 1140

2.96 Iodine 133 1008 2.66

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Astatine 140 920 2.2

Properties of Halogens 1.

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its group number. The charge of a group A nonmetal is the group number minus 8.

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Review Worksheet 1.

Answers will vary. An example is included below: $2\text{H}_2\text{O}(l) + \text{O}_2(g)$ This describes the

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decomposition reaction
of hydrogen peroxide.

Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

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the number of electrons

lost 5. -ite or -ate 6.

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Group 3A (aluminum),

3+' . . .

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