Chapter 9 Mixed Review Stoichiometry

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CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: C3H4(g) + x. O2(g) (3CO2(g) + 2H2O(g) a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C3H4? c. How many moles are in an 8.0 g sample of C3H4? 2. a. What ...

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9-1 Introduction to Stoichiometry pages 275-277 Questions # 1-3. 9-2 Ideal Stoichiometric Calculations pages 280-287 Questions # 1ab,2a,3a . 9-3 Limiting Reactants and Percent Yield pages 288-294 Questions # 1-2 EOC's Page 295 #2,7,10a,12ab,17a,22a,28a,33. Objectives: By the end of this unit you should… Define Stoichiometry.

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Stoichiometry b. Theoretically, how many moles of NH3 will be produced? PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N2 are mixed with 12.0 mol of H2 according to the ...

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left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following equation: N 2(g) 3H 2(g) → 2NH 3(g) N 2; 2.0 mol a.

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percentage yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following equation: N 2(g) 3H 2(g ...

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