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9. If the percentage yield for a chemical reaction is 80.0%, and the theoretical yield is 100 grams, what is the actual yield of product? ... Chapter 9 Stoichiometry Test REVIEW SHEET Last modified by:

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H

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Answer Key Mole/Stoichiometry.Test.Review 1. 6.022×10^{23} particles((atoms,(molecules)) (2. 1mole(=6.022x1023particles((1mole=molar(mass(1mole=22.4L(3. Calculate(the ...

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students...

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Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

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Remember it is a MC test, use the answers ... 3 contains 9×10^{23} oxygen atoms. III. A 200 g sample of CaCO_3 contains 2 moles of CaCO_3 a. I only b. II only c. III only d. I and III only e. I, II, and III Practice Test Ch 3 Stoichiometry Name ____ Per ____ $2\text{MnO}_2 + 4\text{KOH} + \text{O}_2 + \text{Cl}_2 ?$ $2\text{KMnO}_4 + 2\text{KCl} + 2\text{H}_2\text{O}$ 9. For the reaction above, there ...

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stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds.Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the subject of this chapter and it is based on

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