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Questions Answer  
Key:

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Introduction  
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fingerprint is  
unique to each

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person, the color of light emitted by metals heated in a flame is unique to each metal. In this laboratory activity, the characteristic color of light emitted for calcium, copper, lithium,

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potassium,  
sodium, and  
strontium ...  
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questions!!? |  
Yahoo Answers  
If the vinegar  
used in the lab  
had 5.89 g of

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acetic acid per  
100. mL of  
vinegar, what  
volume (in mL)  
of 0.0998 M NaOH  
would you use to  
reach the  
endpoint of the  
titration? ...

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Crystal violet  
is a common,  
beautiful purple

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dye. In strongly basic solutions, the bright color of the dye slowly fades and the solution becomes colorless.

Le Châtelier's  
Principle

So we did a lab  
in class and I  
need help

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finding the  
final answers.  
The data table  
says: the  
initial mass of  
 $\text{CuCl}_2(\text{g})$  was  
.52g Initial  
mass of aluminum  
wire (g) was .74  
The mass of  
"leftover"  
aluminum wire  
(g) was .68 The  
actual mass of



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aluminum reacted  
(g) was .6 The  
actual mass of  
recovered dry  
copper (g) .22 I  
need help  
finding : A) The  
moles of  $\text{CuCl}_2$   
B) Moles of  
aluminum that  
should react ...

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Central — The  
College Board  
Mass of silver  
nitrate- 1.58g  
Mass of copper  
wire- 2.35g Mass  
of empty 100 mL  
beaker- 50.03g  
Mass of leftover  
copper wire-  
2.15g Mass of  
beaker plus  
silver product-  
50.67g 1.

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Calculate mass and moles of copper wire that reacted in this experiment 2.

Calculate the mass and moles of silver metal produced in the reaction 3.

Determine the mole ratio-the ratio of the number of moles

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of silver to ...

Thermodynamics:  
Enthalpy of  
Reaction and  
Hess's Law  
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Calgary, Alberta  
based company  
that has grown  
tremendously  
over the past  
decade. Chemfax  
proudly operates

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out of a 60,000 sq. ft. State-of-the-Art facility which includes one of the largest Class 1, Division 1, flammable liquid handling areas in Canada.

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Bing

- Answer the pre-lab questions that appear at the end of this lab exercise.

The questions should be answered on a separate (new) page of your lab notebook. Be sure to show all work, round

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answers, and include units on all answers. Background information can be found in Chapter 15, especially sections 15.6 your textbook (Brown and LeMay).

6: Types of  
*Page 23/47*

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Chemical  
Reactions  
(Experiment) -  
Chemistry ...  
Thermodynamics:  
Enthalpy of  
Reaction and  
Hess's Law Judy  
Chen Partner:  
Mint Date: 13  
Sept, 2011  
Purpose: The  
purpose of this  
lab is verify



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Hess's law by finding the enthalpies of the reactions; NaOH and HCl, NH<sub>2</sub>Cl and NaOH, and NH<sub>3</sub> and HCl. The overall enthalpy should equal to the sum of enthalpy of the three reactions in

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Lab Techniques  
and Safety by  
Crash Course.  
Lab #1 -  
Analysis of Food  
Dyes in  
Beverages. Lab  
#2 - Percent  
Copper in Brass.  
Lab #3 -

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Gravimetric  
Analysis of  
Calcium in  
Antacid Tablets.  
Lab #4 - Acidity  
of Beverages.  
Lab #5 -  
Separation of a  
Dye Mixture  
Using  
Chromatography.

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Pre-laboratory

Assignment:

Types of

Reactions. Many

of the reactions

use 1 mL of

solution. How

can you estimate

this volume? Do

you need to dry

the test tubes

before using

them for the

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reactions in  
this experiment?  
Why or why not?  
The following  
reactions are  
performed, and  
the results are  
shown below. Use  
these results to  
determine the  
relative ...

AP Chemistry Lab  
/Investigations

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kits, laboratory  
kits, an  
equipment list  
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...

Flame Test Lab  
Questions Answer  
Key: You could  
readily identify  
the elements  
that had obvious  
colors different  
form all the  
others- such as  
copper that gave  
off a blue/green  
color and  
lithium that

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gave off a  
bright red  
color.

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viii

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inquiry

possibilities

for students'

investigations

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see the  
suggestions in  
Using these Labs  
with the AP and  
the IBO Programs  
in this  
Introduction.  
Additionally,  
this manual  
presents teacher-  
developed  
laboratory  
activities using  
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3.) The specific heat of a solution is

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$4.18\text{J}/(\text{g}\cdot\text{C}^\circ)$

and its density is  $1.02\text{g}/\text{mL}$ . The solution is formed by combining  $25.0\text{mL}$  of solution A with  $25.0\text{mL}$  of solution B, with each solution initially at  $21.4^\circ\text{C}$ . The final temperature of the combined

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solutions is

25.3 °C.

Calculate the  
heat of  
reaction,  $q_{\text{rxn}}$ ,  
assuming no heat  
loss to the  
calorimeter.

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modynamics-  
Enthalpy of  
Reaction ...

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revised to  
integrate  
inquiry,  
content, and  
reasoning  
through an  
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