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a state of balance in which the rates of the forward and reverse reactions are equal; no net change in the amount of reactants and products occurs in the chemical system (18.2)  
equilibrium position the relative concentrations of reactants and products of a reaction that has reached equilibrium; indicates whether the reactants or products are favored in the reversible reaction (18.2)

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Modern Chemistry 149 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship between the ion product and the  $K_{sp}$  for that solution, listed on the left. \_\_\_\_\_ The ion product exceeds the  $K_{sp}$ .

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