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*completes the statement or  
answers the question.*

*Chapter 14 - Chemical*

*Kinetics: Part 1 of 17*

*Chemistry I: Chapter 14*

*Test, Gases  $PV=nRT$   $P_1V_1/T_1 =$   
 $P_2V_2/T_2$  760 mm Hg = 1 atmK =*

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$^{\circ}\text{C} + 273$   $1000 \text{ mL} = 1 \text{ L}$   $22.4$   
 $\text{L} = 1 \text{ mol}$  (at STP)  $R = 0.0821$   
 $\text{atm}\cdot\text{L}$   $R = 8.314 \text{ kPa}\cdot\text{L}$   $R = 62.4$   
 $\text{mm Hg}\cdot\text{L}$   $\text{Mol}\cdot\text{K}$   $\text{Mol}\cdot\text{K}$   $\text{Mol}\cdot\text{K}$   
1. Two 500 mL flasks contain  
gas at the same temperature  
and pressure. Flask 1  
contains carbon monoxide,

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CO.

*A.P. Chemistry Practice  
Test: Ch. 14, Acids and  
Bases*

*Chapter 14-Assignment B:  
Molar Volume, Density, and  
Molar Mass Three ratios,*

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*molar volume, density, and molar mass, are important for understanding relationships among the four measurable properties of a gas.*

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*14. A sealed bag of cookies contains 258 mL of air at the store temperature of 25°C. It is left in a hot car where the temperature rises to 39°C. What will the volume be at this temp? 15.*

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*A balloon containing 555 mL of helium at 23°C and pressure is 744 mm Hg is released into the atmosphere.*

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*CHM 2046 Test #3 Review:*

*Chapters 14.8, 14.9, 15, & 16*

*Chapter 14 1. For the*

*following reaction  $K_c =$*

*0.513 at 500 K.  $N_2O_4(g)$*

*?  $2 NO_2(g)$  If a reaction*

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*vessel initially contains an  
N<sub>2</sub>O<sub>4</sub> concentration of  
0.0500 M at 500 K, what are  
the equilibrium  
concentrations of N<sub>2</sub>O<sub>4</sub>  
and NO<sub>2</sub> at 500 K?*

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*Section 14.1: Factors Tha...*

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Child's Lungs have a volume  
of 2.2 L. How many grams of  
air do her lungs hold at a*

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*pressure of 102 kPa and a  
body temperature of 37°C?  
Use a molar mass of 29 g for  
air, which is about 20% O?  
(32 g/mol) and 80% N? (28  
g/mol) .*

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*chemicall... A mixture of tiny particles that are bigger than those in the... A substance being desolved  
Mixture A combination of two or more substances that are not chemicall... Colloids A mixture of tiny particles*

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*that are bigger than those  
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*Volume so the air pressure is greater in your lungs than that of the outside and air rushes out.*

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