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As we noted in Chapter 9.2 the heat released by a reaction carried out at constant volume is identical to the change in internal energy (ΔE) rather than the enthalpy change (ΔH); ΔE is related to ΔH by an expression that

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depends on the change in the number of moles of gas during the reaction.

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Section 9.1 Reactions and Equations pages 282–288
Practice Problems pages 284–287
Write skeleton equations for the following word equations.
1. Hydrogen and bromine gases react to yield hydrogen bromide.

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9.1 Chemical Reactions Are Represented by Chemical ...

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Study Guide Section 9.2 Classifying Chemical Reactions In your textbook, read about synthesis, combustion, decomposition, and replacement reactions. Assume that Q, T, X, and Z are symbols for elements. Match each equation in Column A with the reaction type it represents in Column B.

Column A	Column B
1.	

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Figure 9.2.4 The Enthalpy of Reaction Energy

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changes in chemical reactions are usually measured as changes in enthalpy. (a) If heat flows from a system to its surroundings, the enthalpy of the system decreases, ΔH_{rxn} is negative, and the reaction is exothermic; it is energetically downhill.

Chemical Reactions

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STUDY. Flashcards. Learn. Write. Spell. Test.
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Sophie_Arhart9. Terms in this set (12)
chemical reaction. a change in which one or
more substances are changed into one or more
new substances. reactant. a starting
substance in a chemical reaction.

Chapter 9 Chemical Reactions Pt I

Conceptual Chemistry Chapter 9: How Chemicals

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284 Chapter 9 • Chemical Reactions Figure 9.3 Science, like all other disciplines, has a specialized language that allows specific information to be communicated in a uniform

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manner. This reaction between aluminum and bromine can be

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