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CHAPTER 6 REVIEW Chemical
Bonding SECTION 4 SHORT
ANSWER Answer the following
questions in the space provided.

1. b In metals, the valence
electrons are considered to be (a)
attached to particular positive
ions. (c) immobile. (b) shared by

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all surrounding atoms. (d)
involved in covalent bonds.

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chapter.

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6) purely ionic. b. partially ionic.
c. polar-covalent. d. nonpolar-
covalent. 10. You can estimate

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the degree to which a bond between two atoms is ionic or covalent by calculating the a. distance between the atoms' nuclei. b difference in the atoms' electronegativities. c. atoms' atomic radii. d. number of atoms in the compound.

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Assessment Chapter Test A

Chapter 6 – ModernCHEMICAL
BONDING Chemistry S.Martinez 2.
CHEMICAL BONDING In a
chemical bond between atoms,
their valence electrons are
redistributed in ways that make

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the atoms more stable. The way that atoms are redistributed determines the type of bonding.11/19/2010 S.Martinez 2

Chemical bonding chapter 6 - SlideShare

6. _____ Bonds that possess

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between 5% and 50% ionic character are considered to be (a) ionic. (c) polar covalent. (b) pure covalent. (d) nonpolar covalent.

7. _____ The greater the electronegativity difference between two atoms bonded together, the greater the bond's

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percentage of (a) ionic character.
(c) metallic character. (b)
nonpolar ...

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Practice ...

Covalent bond in which the bonding electrons are equally attra... covalent bond The bond formed when two or more valence electrons are attract... bond length The distance between two bonded atoms at their minimum

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potenti... A mutual electrical attraction between the nuclei and valence... Chemical bonding that results from the electrical attraction b....

Chapter 6 Vocabulary Flashcards

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Objectives: Chapter 6. Notes:
Chapter 6 CH 6 Chemical Bonding
. Reference: CH 6 Common Ion
Table. CH 6 Bond Type Organizer.
CH 6 Dot Notation & Ion
Formation Organizer. CH 6
Drawing Lewis Structures . CH 6
Drawing Lewis Structures

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Explanation. CH 6 Naming Acids.
CH 6 Ionic Bonding Flash. CH 6
Polar Bonding Flash. CH 6
Nonpolar Covalent Flash

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Covalent

Holt Chemistry NY: Chapter 6 - Covalent Compounds. The "magic" element, it is capable of forming single, double, and triple bonds, as well as long chains, rings, and complex shapes, which makes it the most flexible

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element and important element
to organic chemistry.

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Chemical Bonding ... $A + B \rightarrow A^+ + B^-$
Covalent Chemical bond resulting from the sharing of valence electrons between two atoms In a pure covalent the electrons are shared equally between the two atoms (called nonpolar covalent) Most covalent

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result in the electrons being shared unequally between atoms (called polar ...

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covalent bonds and dots adjacent
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