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**The Arrhenius definition of acid and
base is limited to aqueous (that is,**

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water) solutions. Although this is useful because water is a common solvent, it is limited to the relationship between the H^+ ion and the OH^- ion. What would be useful is a more general definition that would be more applicable to other chemical reactions and,

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importantly, independent of H_2O .

pH and pOH CONTINUED -

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pH and pOH CONTINUED

1. 0.01 M HCl $[H^+] = 0.01M$ $pH = -\log(0.01) = 2$

**2. 0.0010 M NaOH $[OH^-] = 0.0010M$
 $pOH = -\log(0.0010) = 3$? $pH = 11$**

3. 0.050 M NaOH

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pH and pOH The pH of a solution indicates how acidic or basic that solution is. pH range of 0-7 acidic 7 neutral 7-14 basic Since ... [H⁺] pH [OH⁻] pOH Acidic or Basic 1. 10⁻⁵ M 5 10⁻⁹ M 9 Acidic 2. 10⁻⁷M 7 10⁻⁷M 7 Neutral 3. 10⁻¹⁰M 10 10⁻⁴ M 4 Basic

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