

## Chemistry M Relationships Answer Key

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Chemistry End of Chapter Exercises. Explain why a sample of pure water at 40 °C is neutral even though  $[H_3O^+] = 1.7 \times 10^{-7} M$ .  $K_w$  is  $2.9 \times 10^{-14}$  at 40 °C. The ionization constant for water ( $K_w$ ) is  $2.9 \times 10^{-14}$  at 40 °C. Calculate  $[H_3O^+]$ ,  $[OH^-]$ , pH, and pOH for pure water at 40 °C.

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Deciding SN1/SN2/E1/E2: The Key Role Of The Alkyl Halide ("Substrate") Having gone through the SN1, the SN2, the E1, and the E2 reactions in turn, we can now say the following: Both substitution reactions and elimination reactions occur with alkyl halides (and related species).

Deciding SN1/SN2/E1/E2 (1) - Master Organic Chemistry

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