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TE - Mr. McKnight
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**10. Dalton's atomic
theory was based on
careful
measurements and
extensive research.**

11. There are no

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instruments

powerful enough to

magnify atoms so

that they can be

seen. The smallest

particle of an

element that retains

the properties of

that 12. element is

called an atom.

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Equations In your**

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***evidence of
chemical reactions.
For each statement,
write yes if evidence
of a chemical
reaction is present.***

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Chapter 10 – The
Mole Section 10.1***

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**Measuring Matter 1.
pair 2. 5 3. dozen 4.
gross 5. 200 6. ream
7. 6,000,000,000 8.
0.5 mol 9. 6.02 1023
10. four moles 11.
6.02 10 Cu atoms 23
1 mol Cu 12. 4 23 4 1
mol CH 6.02 10
molecules CH 13. 23
1 mol Xe 6.02 10
molecules Xe 14. 23
2 2**

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Answers ...**

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**CHAPTER Applying
Scientific Methods**

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CHAPTER

ASSESSMENT A

chemist isolated

four samples, A, B,

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**C, and D. She
obtained the
following atomic
emission spectra of
the samples. 400
500 600 700 29
Nanometers 1.
Examine each
sample's atomic
emission spectra.**

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Solutions Manual

CHAPTER 10

SOLUTIONS

MANUAL 10. Explain

how a mole is

similar to a dozen.

The mole is a unit

for counting 6.02

10^{23} representative

particles. The dozen

is used to count 12

items. 11. Apply

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How does a chemist count the number of particles in a given number of moles of a substance?

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Chemical Reaction.

***The process by
which the atoms of
one or more
substances are
rearranged to form
different
substances; occurs
when there is
change in
temperature, color,
odor, and physical***

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change.
Chapter 10
Answer Key

***Chemistry Matter
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The process by
which the atoms of
one or more
substances are
rearranged to form
different
substances; occurs
when there is
change in***

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**temperature, color,
odor, and physical
change. Key**

chemistry chapter

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The number

6.02×10^{23} , which is

the number of

representative p.... SI

base unit to

measure the amount

of a substance,

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abbreviated.... A formula that specifies the actual number of atoms of each el.... The mass in grams of one mole of any pure substance.

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The Mole chapter of***

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**companion course
helps students learn
the essential
chemistry lessons
of molar mass... for
Teachers for
Schools for Working
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MANUAL The

MoleThe Mole

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Chapter 10 161

Section 10.1

Measuring Matter

page 320-324

Practice Problems

pages 323-324 1.

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Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel.

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and ...

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chemistry chapter
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or more atoms that
covalently bond***

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together to form a unit. The number of carbon atoms in exactly 12 g of pure carbon-12. Two or more atoms that covalently bond together to form a unit. The number of carbon atoms in exactly 12 g of pure carbon-12. Valence Shell Electron Pair Repulsion Theory.

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Chemical Reactions.

BLOCK SCHEDULE

LESSON PLAN 10.

Please note that this

pace is based on

completing selected

sections of the text

in 90 classes,

approximately 90

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***minutes each. Refer
to the Course
Planning Guide on
page xvii of this
booklet for a
complete list of time
allotments assigned
to each section.***

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Vocabulary ...

***The molar mass of a
compound can be
calculated from its
chemical formula
and can be used to
convert from mass***

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to moles of that compound.

Chemical formulas indicate the numbers and types of atoms contained in one unit of the compound. contains one mole of C atoms, two moles of Cl atoms, and two moles of F atoms.

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Chapter 10: The

Mole ...

*The mass in grams
of one mole of any
pure substance*

Avogadro's Number

The number

*6.02×10^{23} , which is
the number of
representative p...*

*Mole SI base unit to
measure the amount
of a substance,*

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abbreviated... The process by which the atoms of one or more substances are r... The starting substance in a chemical equation.

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Mole ...
Step-by-step
solutions to all your***

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