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Chemistry Molecular Approach Nivaldo Tro

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Dipole Moment. When two electrical charges, of opposite sign and equal magnitude, are separated by a distance, an electric dipole is established. The size of a dipole is measured by its dipole moment (μ). Dipole moment is measured in Debye units, which is equal to the distance between the charges multiplied by the charge (1 Debye equals (3.34×10^{-30}) ; C, m).

Dipole Moments - Chemistry LibreTexts

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Chemistry: The Molecular Science (5th Edition) Moore, John W.; Stanitski, Conrad L.

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Quantum numbers. There are four quantum numbers n , l , m_l , and m_s . The principal quantum number n is a positive integer (1,2,3,4) and it represents the energy of the orbital. The angular momentum quantum number l , is from 0 to $n - 1$. The l values of 0, 1, 2, and 3 correspond to the s, p, d and f orbitals, respectively. The magnetic quantum number m_l ranges from $-l$ to $+l$.

1.9A: Ground state electronic Configuration - Chemistry ...

Electron affinity is the amount of energy required to detach one electron from a negatively charged ion of an atom or molecule. It is indicated using the symbol E_a and is usually expressed in units of kJ/mol.

Electron Affinity (Chemistry Glossary Definition)

The electron affinity (E_{ea}) of an atom or molecule is defined as the amount of energy released when an electron is attached to a neutral atom or molecule in the gaseous state to form an anion. $X(g) + e^- \rightarrow X^-(g) + \text{energy}$. Note that this is not the same as the enthalpy change of electron capture ionization, which is defined as negative when energy is released.

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