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of copper-water mixture: We must assume that no energy
will be transferred to or from the beaker containing the
water. Then the magnitude of energy lost by the hot
copper and the energy gained by the cold water will be
equal (but opposite in sign). $q_{\text{copper}} = -q_{\text{water}}$**

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Classify each of the following half-reactions as oxidation

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For the decomposition of Ag_2O : $2 \text{Ag}_2\text{O} (s) \rightarrow 4 \text{Ag} (s)$

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