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QUANTUM NUMBERS WORKSHEET 1.

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State the four quantum numbers, then explain the possible values they may have and what they actually represent. n - Principal Quantum Number: represents the energy level the electron is in, linked to

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With reference to quantum
numbers, explain why the 4f

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sublevel can hold a maximum of 14 electrons. The 4f sublevel has the quantum numbers $n = 4$ and $l = 3$. When $l = 3$, m_l can only be +3, +2, +1, 0, -1, -2, and -3. Thus there are seven orbitals for the 4f

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sublevel. Each orbital can
have electrons

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Quantum Numbers Answers

Numbers. Quantum mechanics is a set of complex mathematics that is used to describe the most probable location of an electron. Shortly after Bohr, a man by the name of Heisenberg proposed an uncertainty

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principle, which stated that
it is impossible to know
both the exact position and
the exact ...

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Table relating quantum numbers to orbital shape. The relationship between three of the four quantum numbers to the orbital shape of simple electronic configuration atoms up through radium (Ra, atomic

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number 88). The fourth quantum number, the spin, is a property of individual electrons within a particular orbital.

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Information: Atoms, Ions,

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**Masses of Subatomic
Particles** The atomic mass
unit (amu) is a special unit
for measuring the mass of
very small particles such as
atoms. The relationship
between amu and grams is the
following: $1.00 \text{ amu} = 1.66 \times$

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10-24g Note the following diagrams of a fluorine atom and a fluorine ion.

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