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Does the reaction depend on the concentration of ClO_2 and the concentration of OH^- - equally In other words, is the rate more dependent on ClO_2 , on OH^- , or is it equally dependent on the concentration of both.

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1. a) How many moles of water are produced when 1.45 moles of propane are combusted? b) How many molecules of water is this? (Remember each mole has 6.02×10^{23})

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molecules.) 2. If 2.35 moles of CO_2 are produced in a reaction, how many moles of H_2O would be produced? 3. Why is this statement false: " If 10 grams of propane burn, you need 50 ...

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Chemistry 2017-18

Here is what a chemical reaction looks like when it is written $\text{Cu}(\text{NO}_3)_2 + \text{Zn} \rightarrow \text{Zn}(\text{NO}_3)_2 + \text{Cu}$ There is a sentence that you can write to describe what happens above: Copper(II) nitrate reacts with zinc metal and yields zinc nitrate and copper metal. Take a

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minute to notice the relationship
between the sentence and the symbols
in the reaction

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are the most involved in chemical

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reactions and bonding. The electron configuration for thallium (#81) is: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^1$. The outermost energy level (not sublevel) is the 6th energy level. How many total electrons does thallium have in the sixth level 3, they are in

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12 moles. For every mole of propane

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that combusts 3 moles of CO₂ are produced, so the number of moles of CO₂ that can be produced when 4 moles of propane combusts = $4(3) = 12$. Verify that if 12.5 moles of propane and 63.2 moles of oxygen were present, then propane is the “ limiting reactant ” and oxygen is

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the excess reactant.

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