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Chemquest 36 Gases And Moles Answer Key

a) How many moles of gas are in his lungs? 0.343 mol b) If

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the diver swims to the surface without exhaling where the temperature is 32°C and the pressure changes to 100.2 kPa, what will the volume of the air in his lungs be? 8.68 L 3. On planet X, 2.78 moles of a gas takes up 1.85 L under a pressure of 74.1 kPa and a temperature of 201°C.

Fig-t

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Gas (# of moles of gas is constant) Gas # of moles of Gas in container (mol) Volume of Container (L) Temperature of Gas (K) Pressure of Gas (kPa) A 0.3008 14.0 672.0 120.0 B 0.3008 16.0 1280 200.0 C 0.3008 10.0 340.0 85.0 Table 2: Experimental Data for Three Different Samples of Gases That Contain Different Amounts of Gas (# of moles of gas is ...

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Intro To Gases Chemquest 35 Answers

ChemQuest 36 Name: Date: Hour Information: Double Replacement Reactions In the previous ChemQuest, you learned about single replacement reactions in which a single atom replaces an ion from another reactant. Study what

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happens in the following reactions. They are called double replacement reactions.

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