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ChemQuest 35 Intro, to chemical Reoctioil\$" Date: Hour: Formula Review 1. Write the formulas for the following compounds: 2. Write the names for the following ...

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118 ChemQuest 39 Name: _____ Date: _____ Hour: _____ Information: Molarity Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little. Critical Thinking Questions 1.

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A solution is made so that it is 2.5 M $\text{Ca}(\text{NO}_3)_2$. Therefore the concentration of Ca^{2+} is 2.5 M and the concentration of NO_3^- is 5.0 M. The total concentration of all particles is 7.5 M. Explain. Since each mole of $\text{Ca}(\text{NO}_3)_2$ breaks up into 1 mole of Ca^{2+} , the concentration of Ca^{2+} would be the same as the concentration of $\text{Ca}(\text{NO}_3)_2$.

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**Calculate the average atomic masses. Round all answers to two decimal places. Record
your answer with the correct units. 1. Iodine is 80% ^{127}I 17% ^{126}I , and 3% ^{128}I .**

**Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of
lithium, which occurs as two isotopes that have the following**

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Chem Quest - Mr. Smith

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How many moles of NaCl will be produced in this reaction? (Remember—you must base

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this answer on how many moles of the limiting reactant that reacted.) 1.5 mol. 0.75 mol of MgCl_2 is the limiting reactant, so we use the 6:3 ratio of NaCl to MgCl_2 and multiply $(0.75)(6/3) = 1.5$ mol.

Chemquest 2 - Intro to Graphs - 9 ChemQuest 2 Name Date ...

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