

## Concept Review Section Changes Of State Answers

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### Concept Review Section Changes Of

Identify two changes of state that require energy to be used, and two changes of state that require the release of energy. Evaporation and sublimation require energy to be used. Condensation and freezing release energy. Apply the law of conservation of mass to explain how mass is conserved when water evaporates.

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Concept Review Section: Changes of State 1. Identify the process that causes each change of state described below.

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### Holt Physics Section Reviews

Concept Review: Energy. 1. energy 2. physical 3. chemical 4. endothermic 5. exothermic 6. kinetic 7. transferred 8. In any chemical or physical change, the total quantity of energy remains constant. Energy cannot be created or destroyed. 9. Heat is the energy transferred between objects that are at different temperatures.

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SECTION: CHANGES OF MATTER 1. a. physical b. chemical c. physical d. physical e. chemical f. chemical 2. The properties of the compound are unaffected. 3. New substances are formed by the application of heat. 4. Physical—properties unaffected, same substances Chemical—new properties, new substances 5. color or smell change, or sound or gas production

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Concept Review j j. 1. A chemical is any substance that has [de fi nit,e, indefinite} composition. Changes in chemicals, or chemical reactions, take place [only in test tubes, all 2. The type and arrangement of [pa.rtibles, crystals] in asample of matter deter mine the properties of the matter.

Concept Review: Changes of State Flashcards | Quizlet

The change in the internal energy of a system is equal to the difference between the internal energy of the final and initial states. According to the law of conservation of energy , energy can be neither created nor destroyed, but it can be transferred to or from the surroundings.

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Name Class Date Concept Review Skills Worksheet. Holt Chemistry 4 Causes of Change Section: Changes in Enthalpy During Chemical Reactions Complete each statement below by writing the correct term in the space provided. In most chemical reactions, the enthalpy change can be measured in terms of energy in the form of released or gained during...

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