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dot Structural formula Space-filling Electron cloud  
Covalent Bonds (pages 165–167) 1. Describe a covalent bond. 2. Circle the letters of molecular models that show orbitals of atoms overlapping when a covalent bond forms.

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The covalent bonding portion starts by showing students how Cl and Cl bond when they come in contact with one another to share unpaired electrons. This is followed by showing oxygen bonding to another oxygen with a double bond with two bonding pairs of electrons. The notes end with a summary of ionic, covalent and metallic bonds.

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For the six that are shared to form the covalent bonds, the phosphorus atom contributed three, and each of the chlorines contributed one. The resulting  $\text{PCl}_3$

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molecule is usually drawn as shown in Figure 6. Figure 6. Three covalent bonds. Each of the three lines represents the shared pair of electrons in a covalent bond.

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they're very stable and because most biological molecules are made with covalent bonds. These biological molecules...

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atoms to form covalent bonds. These bonds are stronger and much more common than ionic bonds in the molecules of living organisms. Covalent bonds are commonly found in carbon-based organic molecules, such as our DNA and proteins.

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covalent bonds. Single Bond.  
when two atoms share one  
pair of electrons. Double  
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When participating in  
covalent bonding, hydrogen  
only needs two electrons to  
have a full valence shell.  
As it has one electron to

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### Answers

start with, it can only make one covalent bond.

Similarly, boron has 3 electrons in its outer shell. This nonmetal typically forms 3 covalent bonds, having a maximum of 6 electrons in its outer shell.

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In covalent bonding, the atoms are unstable because their outer rings of electrons aren't filled up. By sharing electrons with other atoms, these atoms can fill up their outer rings and become stable. In water, for instance, the oxygen atom needs two more electrons to be stable, and the hydrogen atoms each need one.

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