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P3-15 Assuming  
the reactions to  
be elementary: 2  
Anthracene  $\rightarrow$   
Dimer 2 ( )

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Dimer A

Anthracene C C r

Engineering where, K c

= k + /k-

Similarly for

the second

reaction: 3-13

Norbornadiene

Quadricyclane (

) Quadricyclane

Norbornadiene

Norbornadiene C

C r k C K where,

K c = k + /k-



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P3-16 The mistakes are as follows: 1. For any reaction, the rate law cannot be written on the basis of the stoichiometric equation.

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Stoichiometry.  
We shall set up  
Stoichiometric  
Tables using A  
as our basis of  
calculation in  
the following

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reaction. We will use the stoichiometric tables to express the concentration as a function of conversion. We will combine  $C_i = f(X)$  with the appropriate rate law to obtain  $-r_A = f(X)$ .

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