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P3-15 Assuming the reactions to be elementary; 2 Anthracene \rightarrow Dimer 2 () Dimer A Anthracene C C r k C K where, $K_c = k + /k$ -Similarly for the second reaction: 3-13 Norbornadiene Quadricyclane () Quadricyclane Norbornadiene Norbornadiene C C r k C K where, $K_c = k + /k$ -P3-16 The mistakes are as follows: 1. For any reaction, the rate law cannot be written on the basis of the stoichiometric equation.

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Stoichiometry. We shall set up Stoichiometric Tables using A as our basis of calculation in the following reaction. We will use the stoichiometric tables to express the concentration as a function of conversion. We will combine C I = f(X) with the appropriate rate law to obtain -r A = f(X).

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