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chemical equilibrium, the equilibrium shifts to oppose the change until a new equilibrium is reached. 2. Answers may vary. The table salt can be ground up into small particles. The temperature of the water can be increased. 3. a. toward the left b. toward the right c. toward the right d. toward the right 4. a. toward the left b. toward the right

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applied stress is determined. Introduction Chemical Equilibrium No chemical reaction goes to completion. When a reaction stops, some amount of reactants remain.

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Le Châtelier ' s Principle

Concept Review: Measurements and Calculations in Chemistry 1. Accuracy is the extent to which a measurement approaches the true value of a quantity; precision is the extent to which a series of measurements of the same quantity made in the same way agree with one another. 2. Answers may vary. Look at the defense of the answer. 3. a. 7 b ...

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Many chemical reactions do not go to completion but instead attain a state of chemical equilibrium. Chemical equilibrium: A state in which the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant. Equilibrium is a dynamic process \rightleftharpoons the conversions of reactants to products and

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change in a system at equilibrium. •Chemical equilibrium can be disturbed by a stress, but the system soon reaches a new equilibrium. •Le Châtelier ' s principle states that when a system at equilibrium is disturbed, the system adjusts in a way to reduce the change. Section 3 Equilibrium Systems and Chapter 14 Stress

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Skills Worksheet Concept Review

Holt Chemistry 1 Chemical Equilibrium Section: Reversible Reactions and Equilibrium Answer the following questions in the space provided. 1. What is the difference between a reversible reaction and one that goes to completion? 2. When is chemical equilibrium reached? 3. What do the arrows in this equation indicate: $A \rightleftharpoons B$? 4. What is a complex

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