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Isotopes ...  
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1. Here are  
three isotopes  
of an element:*

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$^{12}\text{C}$   $^{14}\text{C}$   $^{13}\text{C}$  c. a.

The element is:

Carbon b. The

number 6 refers

to the Atomic

Number c. The

numbers 12, 13,

and 14 refer to

the Mass Number

d. How many

protons and

neutrons are in

the first

isotope? 6



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protons & 6  
neutrons e. How  
many protons and  
neutrons are in  
the second  
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*Set 1. What is*

*an isotope? 2.*

*Fill in each*

*chart below with*

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information  
about each  
isotope. Assume  
all atoms are  
neutral. 3. Fill  
in the isotope  
names and any  
missing  
information,  
including  
isotope numbers.  
Assume all atoms  
are neutral. 4.  
Uranium-235 and



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*Uranium-238 are  
considered  
isotopes of one*

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*Practice Set 1.*

*What is an  
isotope? 2. What  
does the number  
next to isotopes  
signify? 3. How*



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can you tell

isotopes apart?

For each of the  
following

isotopes, write  
the number of  
protons,  
neutrons, and  
electrons.

Chromium<sup>58</sup>

Chromium<sup>63</sup> # of  
protons 24 24 #

of neutrons 34

39

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*Isotopes 1.*

*Define an*

*isotope. 2. What  
would happen if*

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the number of  
protons were to  
change in an  
atom? 3. Another  
way to write  
isotopes is to  
write the name  
of the element  
then add the  
mass number  
after a dash,  
for example,  $^{14}_6\text{C}$  is  
carbon-14.

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*Isotope*

*composition:*

*Counting*

*protons,*

*electrons, and*

*neutrons. Worked*

*example:*

*Identifying*

*isotopes and*

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*the currently  
selected item.*

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isotope contains  
16 protons, 18  
electrons, and  
16 neutrons.*

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*Practice Set 1.*

*What is an isotope? 2. What does the number next to isotopes signify? 3. How can you tell isotopes apart? For each of the following isotopes, write the number of*



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*protons,  
neutrons, and  
electrons.*

*Chromium-58*

*Chromium-63 # of  
protons # of  
protons # of  
neutrons # of  
neutrons # of  
electrons # of  
electrons*

*Nitrogen-15 ...*

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the number of  
protons,  
electrons, and  
neutrons for  
different  
isotopes If  
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*Chromium-58*

*Chromium-63 # of  
protons # of  
neutrons # of  
electrons*

*Nitrogen-15*

*Nitrogen-20 # of  
protons # of  
neutrons*

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protons is set  
by the atomic  
number. ...*

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