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Chapter 7.3 Slides

Metallic Bonds and Metallic Properties When a metal is subjected to pressure, the metal cations easily slide past one another. If an ionic crystal is struck with a hammer, the blow tends to push the positive ions close together.

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SECTION 7.1 IONS (pages 187–193)

Ionic compounds Metallic bonding Ionic bonds Prentice Hall is a registered trademark of Pearson Education, which is not affiliated with Study.com.

How do you find the number of valence electrons in an atom ...

Ionic and Metallic Bonding 201 ... Metallic Bonds and Metallic Properties helped you learn something new. Metals are made up of closely packed cations rather than neutral atoms. The valence electrons of metal atoms can be modeled as a sea of elec-trons. That is, the valence electrons are mobile and can drift freely from ...

7.3 Bonding in Metals 7 - Henry County School District

IONIC AND METALLIC BONDING Practice Problems In your notebook, answer the following SECTION 7.1 IONS For each element below, state (i) the number of valence electrons in the atom, (ii) the electron dot structure, and (iii) the chemical symbol(s) for the most stable ion. a. Ba 2. How many valence electrons does each of the following atoms have?

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Prentice Hall Chemistry Chapter 7: Ionic and Metallic ...

Metallic Bonds and Metallic Properties 1. Is the following sentence true or false? Metals are made up of cations and valence electrons, not neutral atoms. 2. What are metallic bonds? 3. Name three properties of metals that can be explained by metallic bonding. a. b. c. 4. What happens to an ionic crystal when a force is applied to it? 5.

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7.1 Ions > 23 Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved. Formation of Cations Group 2A Cations Magnesium (atomic number 12) belongs ...

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Chapter 7: Ionic and Metallic Bonding Pearson Chemistry ...

Metallic bonds are the forces of attraction between the free-floating valence electrons and the positively charged metal ions. • These bonds hold metals together.

BONDING AND INTERACTIONS

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Ionic Bonding/Compounds Anions and cations are held together by opposite charges (+ and -) Ionic compounds are called salts. Simplest ratio of elements in an ionic compound is called the formula unit. The bond is formed through the transfer of electrons (lose and gain) Electrons are transferred to achieve noble gas configuration.

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metallic crystal structures. Metallic Bonds and Metallic Properties (pages 201–202) 1. Is the following sentence true or false? Metals are made up of cations, not neutral atoms. _____ 2. What are metallic bonds? 3. Name three properties of metals that can be explained by metallic bonding. a. b. c. 4. What happens to an ionic crystal when a ...

PowerPoint Presentation

In metallic bonds, electrons are shared. They are not free floating, like in ionic or covalent bonds. In metallic bonds, electrons are centralized. Electrons are decentralized in ionic and covalent...

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