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a) 0.15 mol dm⁻³ KOH b) 0.05 mol ...

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1) Use your knowledge of ionic equations to give the molar ratio in which the following acids react with bases. Complete the table to show your answers. Acid Formula of acid Base Formula of base

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TASK 2 – Calorimetry calculations 2

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The equilibrium below was established by allowing 2.00 moles of ethanol and 1.00 moles of ethanoic acid to react at 250C. At equilibrium the mixture contained 0.845 moles of ethyl ethanoate. Calculate Kc at this temperature. $\text{CH}_3\text{COOH}(\text{l}) + \text{CH}_3\text{CH}_2\text{OH}(\text{l}) \rightleftharpoons \text{CH}_3\text{COOCH}_2\text{CH}_3(\text{l}) + \text{H}_2\text{O}(\text{l})$... @ www.CHEMSHEETS.co.uk 26-Jun- 2015 Chemsheets AS 1044 . c) State ...

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-1 k-1. 8 25 cm³ of 2.0 mol dm⁻³ nitric acid was reacted with 25 cm³ of 2.0 mol dm⁻³ potassium hydroxide in an insulated cup. The temperature rose from 20.2°C to 33.9°C.

3.1.6 Chemical equilibria, Le Chatelier's principle and Kc ...

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