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However, F₂ has a lower bond dissociation energy than Cl₂ and Br₂. This is because the atomic size of fluorine is very small. Therefore, the increasing order for bond dissociation enthalpy is: I₂ < F₂ < Br₂ < Cl₂ (2) Bond dissociation energy of a H-X molecule (where X = F, Cl, Br, I) lowers with an increase in the size of an atom.

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